Roll No. ____________

2015 (A)

SSC PART-I (9th CLASS)

CHEMISTRY (NEW SCHEME)

GROUP-II

TIME ALLOWED: 2.45 Hours

SUBJECTIVE

MAXIMUM MARKS: 63

NOTE: - Write same question number and its part number on answer book, as given in the question paper.

SECTION-I

2. Attempt any five parts. 10 = 2 x 5

(i) Define Industrial and Nuclear Chemistry.

(ii) Which gases are present in Air? Write their names.

(iii) Complete the following equations:-

(a) $Cl_2 \xrightarrow{\text{Sunlight}}$ (b) $CH_4 \xrightarrow{\text{Sunlight}}$

(iv) Write down any two properties of Cathode Rays.

(v) Write the Electronic Configuration of Silicon(Si) and Aluminium(Al) atoms.

(vi) How many maximum electrons can be accomodated in L and M shells?

(vii) The atomic radius decreases along a period. Give reason.

(viii) How Newland arranged the elements?

3. Attempt any six parts. 12 = 2 x 6

(i) Define Homogenous Mixture and give one example.

(ii) Metals are good Conductor of Electricity. Why?

(iii) Which are the physical properties of Matter?

(iv) What do you mean by Condensation?

(v) What is meant by Freezing Point?

(vi) Define Aqueous Solution. Also give one example.

(vii) What is the difference between Dilute and Concentrated Solution?

(viii) What is meant by Alloys? Give one example.

(ix) Define Molar solution and give one example.

4. Attempt any five parts. 10 = 2 x 5

(i) Define Oxidation and give an example.

(ii) Determine the Oxidation number of "N" in HNO_3.

(iii) What are Oxidizing and Reducing Agents?

(iv) Define Strong Electrolyte and give one example.

(v) Write one difference between Electrolytic Cell and Galvanic Cell.

(vi) Why the Reactivity of Metals increases down a group?

(vii) Write any two physical properties of Metals.

(viii) Why Electropositivity decreases from left to right in periods?

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SECTION-II

NOTE: - Attempt any three questions.

5.(A) Describe the importance of Chemistry in daily life. 3

(B) Explain the Isotopes of Hydrogen with diagrams. 4

6.(A) Describe Rutherford's Atomic Model. Also write defects in Rutherford's Model. 3

(B) How a Coordinate Covalent Bond is formed? Explain with examples 4

7.(A) Explain three factors affecting the Evaporation. 3

(B) Describe $percentage - \frac{mass}{mass}$ and $percentage - \frac{mass}{volume}$ with the help of examples and formulas. 4

8.(A) Write three methods to prevent Corrosion. 3

(B) Write the reaction of Sodium with H_2O, O_2, Cl_2 and H_2 4

9.(A) We want to prepare 100 cm ^{3} 0.4M of $MgCl_2$. How much $MgCl_2$ is required? 3

(B) Write down four rules for assigning Oxidation number. 4

SECTION-III

10. NOTE: - Attempt any two parts. 5 + 5

(A) (i) Write the material to determine the melting point of Biphenyl. 2

(ii) Write procedure to separate Napthalene from given mixture of Sand and Napthalene by Sublimation. 3

(B) (i) Which material is required for the preparation of 100cm ^{3} Oxalic Acid solution of 0.01M from 0.1M Oxalic Acid solution 2

(ii) Write down the procedure for the preparation of 0.1M NaOH solution having 100cm ^{3} volume. 3

(C) (i) Name the material required to prepare 100 cm ^{3} of 0.01M Hydrochloric Acid solution. 2

(ii) Write down the procedure to prepare Pure Copper Sulphate Crystals from the given Impure Sample of $CuSO_4$. 3

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