

Paper No. 13

Faisalabad Board  
(First Group)ANNUAL  
2016

ACCORDING TO THE NEW PAPER PATTERN OF ALL BOARDS

Roll No.(in Figures): ..... (in Words): .....

Maximum Marks: 12

(OBJECTIVE TYPE)

Time Allowed : 15 Minutes

	A	B	C	D	Write correct option		A	B	C	D	Write correct option		A	B	C	D	Write correct option
1	(A)	(B)	(C)	(D)		5	(A)	(B)	(C)	(D)		9	(A)	(B)	(C)	(D)	
2	(A)	(B)	(C)	(D)		6	(A)	(B)	(C)	(D)		10	(A)	(B)	(C)	(D)	
3	(A)	(B)	(C)	(D)		7	(A)	(B)	(C)	(D)		11	(A)	(B)	(C)	(D)	
4	(A)	(B)	(C)	(D)		8	(A)	(B)	(C)	(D)		12	(A)	(B)	(C)	(D)	

NOTE: Four possible answers A, B, C and D to each question are given. The choice which you think is correct, fill that circle in front of that question with Marker or Pen ink. Cutting or filling two or more circles will result in zero mark in that question.

Q1.

12

1. Transfer of electrons between atoms results in:

- (A) metallic bonding (B) ionic bonding  
(C) covalent bonding (D) coordinate covalent bonding

2. One atmospheric pressure is equal to how many Pascals?

- (A) 101325 (B) 10325 (C) 106075 (D) 10523

3. Which is heterogeneous mixture?

- (A) milk (B) ink (C) milk of magnesia (D) sugar solution

4. The most common example of corrosions:

- (A) chemical decay (B) rusting of iron (C) rusting of aluminum (D) rusting of tin

5. The oxidation number of chromium in  $K_2Cr_2O_7$ :

- (A) +2 (B) +6 (C) +7 (D) +14

6. Which metal burns with a brick red flame?

- (A) sodium (B) magnesium (C) iron (D) calcium

7. The most abundant element occurring in the ocean is:

- (A) oxygen (B) hydrogen (C) nitrogen (D) silicon

8. Which molecule is not triatomic?

- (A)  $H_2$  (B)  $O_3$  (C)  $H_2O$  (D)  $CO_2$

9. Who presented the concept of orbit in atom?

- (A) J.J. Thomson (B) Rutherford (C) Bohr (D) Plancks

10. Along the period which one decreases?

- (A) atomic radius (B) ionization energy (C) electron affinity (D) electronegativity

11. Which halogen has the lowest electronegativity?

- (A) fluorine (B) chlorine (C) bromine (D) iodine

12. A bond formed between two non metals is expected to be:

- (A) covalent (B) ionic (C) coordinate covalent (D) metallic