Faisalabad Board 2016 (First Group)

Roll No.(in Figures): .... (in Words): ....

Maximum Marks: 48 **(SUBJECTIVE TYPE)** Time Allowed :1.45 Hours

PART - I

Q2. Write short answers to any FIVE (5) questions: (5×2=10)

(i) Define biochemistry and environmental chemistry.

(ii) Write two differences between molecule and molecular ion.

(iii) Write down the isotopes of chlorine.

(iv) Write down the two defects in Rutherford's model.

(v) An element has 5 electrons in M shell. Find out its atomic number.

(vi) Define period. How many periods are there in long form of periodic table?

(vii) Show why in a period size of atom decreases when we move from left to right in period?

(viii) Define ionization energy.

Q3. Write short answers to any FIVE (5) questions: (5×2=10)

(i) Define intermolecular force and give example.

(ii) What are dipole-dipole forces?

(iii) Define hydrogen bonding and give one example.

(iv) Why ice floats on surface of water?

(v) Define Charles's law. Also write its mathematical form.

(vi) Why evaporation causes cooling?

(vii) What is the effect of surface area on evaporation?

(viii) What is the effect of external pressure on boiling point?

Q4. Write short answers to any FIVE (5) questions: (5×2=10)

(i) Why the suspension does not form a homogeneous mixture?

(ii) Define tyndall effect and give one example.

(iii) What do you mean by weak electrolyte? Give one example.

(iv) Define oxidation and reduction.

(v) What is an alloy? Give one example.

(vi) Why reactivity of metals increases down the group?

(vii) Define amorphous solids with one example.

(viii) Write the names of two allotropic forms of sulphur.

PART - II

Note: Attempt any TWO questions. (9×2=18)

Q5. (a) Explain any five important branches of Chemistry. 5

(b) Write notes on isotopes of carbon and chlorine. 4

Q6. (a) Describe the five physical properties of metals. 5

(b) Define vapour pressure. Explain the three factors which affect vapour pressure. 4

Q7. (a) Describe colloid and suspension with examples. 5

(b) Write rules for assigning oxidation state. 4