

# Faisalabad Board 2016 (First Group)

Roll No.(in Figures): ..... (in Words): .....

Maximum Marks: 48 (SUBJECTIVE TYPE) Time Allowed :1.45 Hours

## PART - I

**Q2. Write short answers to any FIVE (5) questions:** (5×2=10)

- (i) Define biochemistry and environmental chemistry.
- (ii) Write two differences between molecule and molecular ion.
- (iii) Write down the isotopes of chlorine.
- (iv) Write down the two defects in Rutherford's model.
- (v) An element has 5 electrons in M shell. Find out its atomic number.
- (vi) Define period. How many periods are there in long form of periodic table?
- (vii) Show why in a period size of atom decreases when we move from left to right in period?
- (viii) Define ionization energy.

**Q3. Write short answers to any FIVE (5) questions:** (5×2=10)

- (i) Define intermolecular force and give example.
- (ii) What are dipole-dipole forces?
- (iii) Define hydrogen bonding and give one example.
- (iv) Why ice floats on surface of water?
- (v) Define Charles's law. Also write its mathematical form.
- (vi) Why evaporation causes cooling?
- (vii) What is the effect of surface area on evaporation?
- (viii) What is the effect of external pressure on boiling point?

**Q4. Write short answers to any FIVE (5) questions:** (5×2=10)

- (i) Why the suspension does not form a homogeneous mixture?
- (ii) Define tyndall effect and give one example.
- (iii) What do you mean by weak electrolyte? Give one example.
- (iv) Define oxidation and reduction.
- (v) What is an alloy? Give one example.
- (vi) Why reactivity of metals increases down the group?
- (vii) Define amorphous solids with one example.
- (viii) Write the names of two allotropic forms of sulphur.

## PART - II

**Note: Attempt any TWO questions.** (9×2=18)

**Q5. (a) Explain any five important branches of Chemistry.**

**(b) Write notes on isotopes of carbon and chlorine.**

**Q6. (a) Describe the five physical properties of metals.**

**(b) Define vapour pressure. Explain the three factors which affect vapour pressure.**

**Q7. (a) Describe colloid and suspension with examples.**

**(b) Write rules for assigning oxidation state.**