

Gujranwala Board 2016 (Second Group)

Roll No.(in Figures):

(in Words):

Maximum Marks: 48

(SUBJECTIVE TYPE)

Time Allowed :1.45 Hours

PART - I

Q2. Write short answers to any FIVE (5) questions:

(5×2=10)

- Define Nuclear Chemistry.
- What is meant by mole?
- Define isotopes. Give an example.
- Write names of two elements which are produced in U-235 fission reaction.
- Define electronic configuration.
- State Newlands Octaves Law.
- Write names of any four elements or symbols of seventeen group of periodic table.
- Define Shielding Effect.

Q3. Write short answers to any FIVE (5) questions:

(5×2=10)

- Which type of covalent bond is formed in nitrogen (N_2) gas?
- Define metallic bond.
- Define double covalent bond and triple covalent bond.
- Differentiate between "lone pair" and "bond pair" of electrons.
- Write two properties of liquid state of matter.
- Write names of the factors which affect evaporation.
- Why are gases compressible?
- Define boiling point.

Q4. Write short answers to any FIVE (5) questions:

(5×2=10)

- What is an alloy? Give one example.
- How is molar solution prepared?
- Differentiate between oxidizing and reducing agents.
- Why is galvanizing done?
- Define electroplating.
- Define electropositivity.
- Write two uses of sodium metal.
- Write two differences between alkali metals and alkaline earth metals.

PART - II

Note: Attempt any TWO questions.

(9×2=18)

Q5. (a) Write five differences between compound and mixture.

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(b) Write postulates of Bohr's Atomic Theory.

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Q6. (a) Explain boiling point with the factors which affect it.

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(b) Narrate four properties of ionic compounds.

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Q7. (a) Explain the effect of temperature on solubility of different salts.

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(b) Describe the working of an electrolytic cell.

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