

Roll No.(in Figures): (in Words):

Maximum Marks: 48

(SUBJECTIVE TYPE)

Time Allowed :1.45 Hours

PART - I

Q2. Write short answers to any FIVE (5) questions: (5×2=10)

- (i) Define Chemical Formula with examples.
- (ii) Differentiate between Atom and Ion.
- (iii) Write two properties of Canal Rays.
- (iv) Write the Electronic Configuration of Argon (At No = 18) and Boron (At No = 5).
- (v) Define Isotopes. Give examples.
- (vi) What is meant by Electronegativity?
- (vii) Briefly describe the trend of Atomic Radius in a group.
- (viii) Define Period and Group.

Q3. Write short answers to any FIVE (5) questions: (5×2=10)

- (i) Why a Dipole develops in a Molecule?
- (ii) What are the Van der Waals Forces?
- (iii) What is a Dative Covalent Bond?
- (iv) Why does Ice float over water?
- (v) Convert-30°C to Kelvin Temperature.
- (vi) What is Diffusion? Give one example
- (vii) Differentiate between Evaporation and Boiling Point.
- (viii) Define Boyle's Law of Gases.

Q4. Write short answers to any FIVE (5) questions: (5×2=10)

- (i) What do you mean by "Like dissolves like"?
- (ii) Why do we stir paints thoroughly before using?
- (iii) Differentiate between Strong and Weak Electrolytes.
- (iv) Why Steel is plated with Nickel electroplating before the electroplating of Chromium?
- (v) Why an Iron grill is painted frequently?
- (vi) Why Copper is used for making Electrical Wires?
- (vii) Why Magnesium is harder than Sodium?
- (viii) Why does Electropositivity increase from top to bottom in a group?

PART - II

Note: Attempt any TWO questions. (9×2=18)

- Q5. (a) Explain types of Molecules in detail. 5
- (b) Explain results of the experiment of Rutherford's Atomic Model. 4
- Q6. (a) Define Ionic Compound with an example. Give its properties. 5
- (b) Describe Solid state of Matter. Explain its various properties. 4
- Q7. (a) Define Molarity. Write its unit and write its formula to prepare Molar solution. 5
- (b) Discuss the construction and working of the cell in which electricity is produced. 4