Multan Board 2017 (Second Group)

Roll No.(in Figures): ..... In Words): .....

Maximum Marks: 48 **(SUBJECTIVE TYPE)** Time Allowed :1.45 Hours

PART - I

Q2. Write short answers to any FIVE (5) questions: (5×2=10)

(i) State Nuclear Chemistry and Analytical Chemistry.

(ii) Define Valency with an examlpe.

(iii) Write two differences of Bohr's Atomic Theory and Rutherford's Atomic Theory.

(iv) Write any two properties of the Positive Rays.

(v) What is the nature of Charge on Cathode Rays? Also write its origion.

(vi) What is the trend of Ionization Energy in a Period?

(vii) Define Electronegativity. Also write Electronegativity of Chlorine.

(viii) State Shielding Effect. Write its trend in period.

Q3. Write short answers to any FIVE (5) questions: (5×2=10)

(i) What is the difference between Duplet and Octet Rule?

(ii) Why Boiling Point of Water is higher than that of Alcohol?

(iii) Why Noble Gases are Non-reactive?

(iv) Write down two properties of Non-polar Compounds.

(v) Define Diffusion and give an example.

(vi) Write down the names of Allotropic forms of Phosphorus and Sulphur.

(vii) Define Boyle's Law.

(viii) Why Evaporation causes Cooling?

Q4. Write short answers to any FIVE (5) questions: (5×2=10)

(i) Define Electrolysis.

(ii) Define Tyndall Effect.

(iii) Differentiate between Oxidizing Agents and Reducing Agents.

(iv) Briefly describe the process of Tin Coating.

(v) Find out the Oxidation number of Sulphur in Na_2SO_4.

(vi) Why Metallic Character decreases along a period from left to right in a pariodic Table?

(vii) How will you compare the Electropositivity of Alkali Metals and Alkaline Earth Metals?

(viii) Write chemical reaction of Sodium with Oxygen.

PART - II

Note: Attempt any TWO questions. (9×2=18)

Q5. (a) Explain types of Molecules with examples. 5

(b) Describe the Isotopes of Hydrogen and Carbon with diagram. 4

Q6. (a) State and explain Coordinate Covalent Bond with examples. 5

(b) What is Vapour pressure? How it is affected by different factors? 4

Q7. (a) Define Solubility and explain general principle of Solubility. 5

(b) Write down the rules for assigning the Oxidation state. 4