

**PAPER NO.**  
**40**

**RAWALPINDI**  
**BOARD**  
SECOND GROUP

**ANNUAL**  
**2018**

ACCORDING TO THE NEW PAPER PATTERN OF ALL BOARDS

Roll No.(in Figures): ..... (in Words): .....

Maximum Marks: 12

**OBJECTIVE TYPE**

Time Allowed : 15 Minutes

	A	B	C	D	Write correct option		A	B	C	D	Write correct option		A	B	C	D	Write correct option
1	(A)	(B)	(C)	(D)		5	(A)	(B)	(C)	(D)		9	(A)	(B)	(C)	(D)	
2	(A)	(B)	(C)	(D)		6	(A)	(B)	(C)	(D)		10	(A)	(B)	(C)	(D)	
3	(A)	(B)	(C)	(D)		7	(A)	(B)	(C)	(D)		11	(A)	(B)	(C)	(D)	
4	(A)	(B)	(C)	(D)		8	(A)	(B)	(C)	(D)		12	(A)	(B)	(C)	(D)	

**NOTE:** Four possible answers A, B, C and D to each question are given. The choice which you think is correct, fill that circle in front of that question with Marker or Pen ink. Cutting or filling two or more circles will result in zero mark in that question.

Q1.

12

1. Number of elements in normal periods is:

- (A) 18 (B) 08 (C) 32 (D) 10

2. Law of octaves was put forwarded by:

- (A) Newlands (B) Dobereiner (C) Mendeleev (D) Moseley

3. How many electrons are involved in the formation of double covalent bond?

- (A) 3 (B) 1 (C) 2 (D) 4

4. Which one of the following is amorphous solid?

- (A) diamond (B) sodium chloride (C) potassium chloride (D) plastic

5. An example of universal solvent is:

- (A) benzene (B) water (C) alcohol (D) ether

6. The number of grams of solute dissolved in 100 grams of solution, the percentage is called:

- (A) percentage  $\frac{\text{mass}}{\text{mass}}$  (B) percentage  $\frac{\text{mass}}{\text{volume}}$  (C) percentage  $\frac{\text{volume}}{\text{mass}}$  (D) percentage  $\frac{\text{volume}}{\text{volume}}$

7. Pure water is an example of:

- (A) weak electrolyte (B) strong electrolyte (C) non-electrolyte (D) strong acid

8. Anode of down's cell is made up of:

- (A) iron (B) steel (C) graphite (D) zinc

9. A metal which is soft and can be cut with knife is:

- (A) calcium (B) sodium (C) magnesium (D) iron

10. An element which occurs in gaseous state is?

- (A) mercury (B) gold (C) oxygen (D) sodium

11. Gram atomic mass of hydrogen is:

- (A) 1.08 g (B) 1.008 amu (C) 2.016 g (D) 1.008 g

12. When uranium -235 breaks up, it produces:

- (A) neutron (B) electron (C) proton (D) none