

Roll Number
In Figures:- _____
In Words:- _____

PR IX(01)15
CHEMISTRY (New)
9th (Fresh/Reappear)
Fic. No. _____

Superintendent
Signature/Stamp: _____

(For Board's Office use only)

CHEMISTRY (New)

9th (Fresh/Reappear)

Fic. No. _____

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Time Allowed: 3 Hours

Marks: 65

Note:- There are THREE section in this paper i.e Section A,B, and C.

Attempt Section-A on the same paper and return it to the Superintendent within the given time.

No marks will be awarded for Cutting, Erasing or Overwriting. Marks of Identification will lead to UFM case, Mobile Phoe etc. are not allowed in the examination hall.

Section-A

Time Allowed:15 minutes SECTION-A Marks:- 12

Q-1. Write the correct option i.e. A, B, C or D in the empty box provided opposite to each part.

i. Which one of the given periods is shortes period?

(A) Second (B) F i r s t (C) T h i r d (D) None B
period Period Period

ii. What geometrical shape NH₃ has?

(A) Linear (B) Triangular (C) Tetrahedral (D) Pyramidal

iii. Which one of the given is an electrolyte?

(A) sugar (B) Table salt (C) Plastic (D) None of B
solution solution these

iv. S-block elements are _____.

(A) Non (B) Metals (C) Metalloid (D) Transition B
metals

v. The fundamental unit of an elements is _____.

(A) Raidcal (B) Atom (C) Molecule (D) None B

vi. The K-shell can accommodate maximum number of electrons.

(A) 2 (B) 6 (C) 8 (D) 18

vii. The particles of salute in colloids are _____.

(A) Visible (B) Invisible (C) In size like (D) In size like A
suspension solution
solute solute

viii. The halogens are _____.

(A) Inert (B) Reactive (C) electropositive (D) Metallic B

ix. Which one of the given molecules is diatomic?

(A) H₂O (B) H₂O₂ (C) SO₂ (D) HCl D

x. Uranium has how many isotopoes?

(A) 3 (B) 2 (C) 4 (D) None

xi. At higher altitude water boils at _____.

(A) Any (B) 100^oC (C) Higher (D) Lower D
temp then 100^oC then 100^oC

xii. Which of the given compounds has same empirical and molecular formula?

(A) C₆H₆ (B) H₂O₂ (C) H₂O (D) C₆H₁₂O₆

For English Medium Students

PR IX (01) 15
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Note:- Time allowed for Section-B and Section-C is 2 Hours and 45 Minutes.

SECTION-B Marks: 32

Q-II Answer any EIGHT parts. Each part carries equal marks.

- What is Molecule? Also describe the types of molecule.
- Find out the number of protons, electrons and neutrons in the given.

$^{107}_{47}\text{Ag}$	$^{23}_{11}\text{Na}$	$^{56}_{26}\text{Fe}$	$^{207}_{82}\text{Pb}$
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- Describe the major objections raised against Ruther Ford Atomic Model.
- Draw the structures of Hydrogen isotopes.
- Define shielding effect. Does is vary in a period?
- Why do atoms form chemical bond?
- Write down four important properties of covalent bond.
- Differentiate between Diffusion and Effusion.
- Give reasons that a gas expands on heating.
- What is Molarity? Explain briefly.
- What are weak and strong electrolytes?

SECTION-C Marks: 21

Note:- Attempt any THREE questions. All questions carry equal marks.

Q-III (a) Define chemistry. Explain any four braches of Chemistry.

(b) Why atom is a neutral partick?

Q-IV (a) Define electronegativity of an element. Discuss its periodic variation in the periodic table.

(b) Define covalent bond and explain its types.

Q-V (a) What is the difference between avaporation and boiling?

(b) What is Solubility? Explain the factors which affect solubility.

Q-VI (a) What is Electrolytic Cell?

(b) Write down the uses of noble metals.