

## Bahawalpur Board 2018 (First Group)

Roll No.(in Figures): ..... (in Words): .....

Maximum Marks: 48 **SUBJECTIVE TYPE** Time Allowed :1.45 Hours**(PART - I)**

Q2. Write short answers to any FIVE (5) questions: (5×2=10)

- Write two application of nuclear chemistry.
- Define molecular formula and give one example.
- What is the difference between atom and ion?
- How U-235 is used for power generation?
- What are the defects of Rutherford's atomic model?
- Why the ionization energy increases in period?
- Why the elements are called 'S' and 'P' block elements?
- What are the contributions of Moseley in periodic table?

Q3. Write short answers to any FIVE (5) questions: (5×2=10)

- How an atom can accommodate 08 electrons in its valence shell. Write two ways.
- What are inter molecular forces?
- Define malleable and ductile property of metals.
- What is standard atmospheric pressure?
- Why densities of gases are low as compared to liquids?
- Define unsaturated solution.
- Explain (% m/v) mass/volume solution with one example.
- What is Tyndall effect?

Q4. Write short answers to any FIVE (5) questions: (5×2=10)

- Define reducing agent and also give an example.
- Why is steel plated with nickel before the electroplating of chromium?
- How can you prove with an example that conversion of atom to an ion is a reduction reaction?
- Calculate the oxidation number of sulphur in  $H_2SO_4$ . As O.N. of H = +1 and O.N. of O = -2
- Write the names of any two moderately reactive metals.
- How oxygen react with magnesium?
- State two uses of sodium.
- Why is sodium metal more reactive than magnesium metal?

**(PART - II)**

Note: Attempt any TWO questions. (2×9=18)

Q5. (a) State postulates of Bohr's atomic theory. 5

(b) State any four types of molecules. 4

Q6. (a) Define hydrogen bonding. Explain that how these forces affects the physical properties of compound? 5

(b) What is evaporation? Discuss the factors affecting evaporation. 4

Q7. (a) Describe the electroplating of chromium in detail. 5

(b) Write down any four characteristics of solution. 4