

Time: 2 Hours 40 Minutes

SECTION-B

Marks: 32

2. Attempt any eight of the following. All carry equal marks.

- i. Differentiate between elements and compounds with examples.
- ii. Find out the molecular mass of C_6H_6 (atomic mass: C = 12, H = 01)
- iii. Find out the electronic configuration of: Na^{11} and Cl^{17}
- iv. Write down four uses of isotopes.
- v. Why S-block elements have two groups only?
- vi. Why do atoms form chemical bonds?
- vii. Give four physical properties of ionic compounds.
- viii. Why gases are compressible but solids are not? Give reason.
- ix. Differentiate between solution and suspension.
- x. What are strong and weak electrolytes? Support your answer with examples.
- xi. Write four uses of sodium metal.

SECTION-C

Marks: 21

NOTE: Attempt any three of the following questions. All questions carry equal marks.

3.
 - i. What do you understand by the terms mole and avogadro's number? Explain these with suitable examples.
 - ii. Discuss the isotopes of hydrogen.
4.
 - i. Explain electron affinity and its trends in modern periodic table.
 - ii. Discuss dipole-dipole interaction with least one example.
5.
 - i. What are solids? Differentiate between amorphous and crystalline solids.
 - ii. Explain any three factors affecting solubility.
6.
 - i. Elaborate the difference between oxidation and reduction reactions with examples.
 - ii. What is corrosion? Explain rusting of iron as an example.